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Name _____

Naming Ionic Compounds Practice Worksheet

Name the following ionic compounds:

- 1) NH_4Cl _____
- 2) $\text{Fe}(\text{NO}_3)_3$ _____
- 3) TiBr_3 _____
- 4) Cu_3P _____
- 5) SnSe_2 _____
- 6) GaAs _____
- 7) $\text{Pb}(\text{SO}_4)_2$ _____
- 8) $\text{Be}(\text{HCO}_3)_2$ _____
- 9) $\text{Mn}_2(\text{SO}_3)_3$ _____
- 10) $\text{Al}(\text{CN})_3$ _____

Write the formulas for the following compounds:

- 11) chromium (VI) phosphate _____
- 12) vanadium (IV) carbonate _____
- 13) tin (II) nitrite _____
- 14) cobalt (III) oxide _____
- 15) titanium (II) acetate _____
- 16) vanadium (V) sulfide _____
- 17) chromium (III) hydroxide _____
- 18) lithium iodide _____
- 19) lead (II) nitride _____
- 20) silver bromide _____

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Naming Binary Compounds

P_2O_5 = **diphosphorus pentoxide**

CO_2 = **carbon dioxide**

CO = **carbon monoxide**

N_2O = **dinitrogen monoxide**

Formulas

- Molecular CH_4O
- Structural
- Condensed structural CH_3OH
- Wedge and dash
- Ball and stick model
- Space filling model

Name: KEY Block: _____

Assignment #1 – Compound Names and Formulas Single-valent ions only

A. Name these Compounds

- | | |
|---|--|
| 1. Li_2S <u>Lithium Sulfide</u> | 10. GeF_4 <u>Germanium Fluoride</u> |
| 2. CaO <u>Calcium Oxide</u> | 11. Ga_2O_3 <u>Gallium Oxide</u> |
| 3. NaF <u>Sodium Fluoride</u> | 12. EsCl_6 <u>Einsteinium Chloride</u> |
| 4. CaBr_2 <u>Calcium Bromide</u> | 13. Fr_2O_3 <u>Francium Oxide</u> |
| 5. MgCl_2 <u>Magnesium Chloride</u> | 14. Mg_3N_2 <u>Magnesium Nitride</u> |
| 6. BBr_3 <u>Boron Bromide</u> | 15. Rb_2O <u>Rubidium Oxide</u> |
| 7. Cs_2O <u>Cesium Oxide</u> | 16. RaO <u>Radium Oxide</u> |
| 8. FrBr <u>Francium Bromide</u> | 17. SrO <u>Strontium Oxide</u> |
| 9. Ag_2S <u>Silver Sulfide</u> | 18. Tc_2O_7 <u>Technetium Oxide</u> |

B. Write the correct chemical formula for these compounds by balancing the ionic charges

- | | |
|--|---|
| 1. sodium chloride <u>NaCl</u> | 11. hydrogen oxide <u>H_2O</u> |
| 2. magnesium fluoride <u>MgF_2</u> | 12. francium nitride <u>Fr_3N</u> |
| 3. silver oxide <u>Ag_2O</u> | 13. rubidium phosphide <u>Rb_3P</u> |
| 4. indium bromide <u>InBr_3</u> | 14. potassium oxide <u>K_2O</u> |
| 5. zinc bromide <u>ZnBr_2</u> | 15. beryllium sulphide <u>BeS</u> |
| 6. neodymium oxide <u>Nd_2O_3</u> | 16. lithium sulphide <u>Li_2S</u> |
| 7. thorium sulphide <u>ThS_2</u> | 17. hydrogen bromide <u>HBr</u> |
| 8. actinium oxide <u>Ac_2O_3</u> | 18. strontium nitride <u>Sr_3N_2</u> |
| 9. radium bromide <u>RaBr_2</u> | 19. calcium oxide <u>CaO</u> |
| 10. cesium oxide <u>Cs_2O</u> | 20. tantalum nitride <u>Ta_3N_5</u> |

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How to name molecular compounds examples. Mixed ionic/covalent compound naming answers. How to name ionic compounds. More mixed ionic and covalent compound naming answers. Mixed ionic covalent and acid compound naming and formula writing worksheet answers.

What is the name of the covalent compound CCl_4 ? It's carbon tetrachloride. Carbon tetrachloride is an important nonpolar covalent compound. You determine its name based on the atoms present in the compound. By convention, the positively-charged (cation) part of the molecule is named first, followed by the negatively-charged (anion) part. The first atom is C, which is the element symbol for carbon. The second part of the molecule is Cl, which is the element symbol for chlorine. When chlorine is an anion, it is called chloride. There are 4 chloride atoms, so the name for 4, tetra, is used. This makes the molecule's name carbon tetrachloride. CCl_4 goes by many names besides carbon tetrachloride, including tetrachloromethane (IUPAC name), carbon tet, Halon-104, benziform, Freon-10, methane tetrachloride, Tetrasol, and perchloromethane. It is an organic compound that is a colorless liquid with a distinctive sweet odor, resembling that of ether or tetrachloroethylene used by dry cleaners. It's used primarily as a refrigerant and as a solvent. As a solvent, it is used to dissolve iodine, fats, oils, and other nonpolar compounds. The compound has also been used as a pesticide and fire extinguisher. Although carbon tetrachloride used to be widely available and used, it has been replaced by safer alternatives. CCl_4 is known to cause liver failure. It also damages the nervous system and kidneys and may cause cancer. Primary exposure is via inhalation. Carbon tetrachloride is a greenhouse gas known to cause ozone depletion. In the atmosphere, the compound has an estimated lifetime of 85 years. Ionic compounds consist of cations (positive ions) and anions (negative ions). Ionic compound nomenclature or naming is based on the names of the component ions. In all cases, ionic compound naming gives the positively charged cation first, followed by the negatively charged anion. Here are the principal naming conventions for ionic compounds, along with examples to show how they are used: A Roman numeral in parentheses, followed by the name of the element, is used for elements that can form more than one positive ion. There is no space between the element name and the parentheses. This notation is usually seen with metals since they commonly display more than one oxidation state or valence. You can use a chart to see the possible valences for the elements. Fe^{2+} Iron(II) Fe^{3+} Iron(III) Cu^+ Copper(I) Cu^{2+} Copper(II) Example: Fe_2O_3 is iron(III) oxide. Although Roman numerals are used to denote the ionic charge of cations, it is still common to see and use the endings -ous or -ic. These endings are added to the Latin name of the element (e.g., stannous/stannic for tin) to represent the ions with lesser or greater charge, respectively. The Roman numeral naming convention has wider appeal because many ions have more than two valences. Fe^{2+} Ferrous Fe^{3+} Ferric Cu^+ Cuprous Cu^{2+} Cupric Example: FeCl_3 is ferric chloride or iron(III) chloride. The -ide ending is added to the name of a monatomic ion of an element. H- Hydride F- Fluoride O²⁻ Oxide S²⁻ Sulfide N³⁻ Nitride P³⁻ Phosphide Example: Cu_3P is copper phosphide or copper(I) phosphide. Some polyatomic anions contain oxygen. These anions are called oxyanions. When an element forms two oxyanions, the one with less oxygen is given a name ending in -ite and the one with more oxygen are given a name that ends in -ate. NO₂- Nitrite NO₃- Nitrate SO₃²⁻ Sulfite SO₄²⁻ Sulfate Example: KNO_2 is potassium nitrite, while KNO_3 is potassium nitrate. In the case where there is a series of four oxyanions, the hypo- and per- prefixes are used in conjunction with the -ite and -ate suffixes. The hypo- and per- prefixes indicate less oxygen and more oxygen, respectively. ClO- Hypochlorite ClO₂- Chlorite ClO₃- Chlorate ClO₄- Perchlorate Example: The bleaching agent sodium hypochlorite is NaClO. It is also sometimes called the sodium salt of hypochlorous acid. Polyatomic anions sometimes gain one or more H⁺ ions to form anions of a lower charge. These ions are named by adding the word hydrogen or dihydrogen in front of the name of the anion. It is still common to see and use the older

naming convention in which the prefix bi- is used to indicate the addition of a single hydrogen ion. HCO3- Hydrogen carbonate or bicarbonateHSO4- Hydrogen sulfate or bisulfateH2PO4- Dihydrogen phosphate Example: The classic example is the chemical name for water, H2O, which is dihydrogen monoxide or dihydrogen oxide. Dihydrogen dioxide, H2O2, is more commonly called hydrogen dioxide or hydrogen peroxide.

Get 24/7 customer support help when you place a homework help service order with us. We will guide you on how to place your essay help, proofreading and editing your draft – fixing the grammar, spelling, or formatting of your paper easily and cheaply. Inorganic Chemistry 4th edition, Catherine Housecroft Formal theory. Formally, a string is a finite, ordered sequence of characters such as letters, digits or spaces. The empty string is the special case where the sequence has length zero, so there are no symbols in the string. A mixture (is/is not) a chemical combining of substances. Changes: Physical or Chemical? - Answer Key 1 When atoms gain, lose, or share electrons 2 Four 3 True 4 False 5 Physical changes 6 All of the above 7 Change in shape 8 Photosynthesis Return to edHelper. The compound A 1 2 S 3 is an ionic compound. substance can be easily observed. Description AP4238 Science Laboratory Safety Test, Pad Download Free Flinn Scientific Ionic Formula Writing Kit Answers practical emphasis, hugely popular in the first edition, features applications in the upper crust, including petroleum and groundwater geology, highlighting the importance of structural geology in exploration and exploitation of petroleum and water ... A metalloid is a type of chemical element which has a preponderance of properties in between, or that are a mixture of, those of metals and nonmetals.There is no standard definition of a metalloid and no complete agreement on which elements are metalloids. Despite the lack of specificity, the term remains in use in the literature of chemistry.. The six commonly recognised metalloids are ... Ionic bond: An ionic bond forms between two atoms when an electron is transferred from one atom to the other, forming a positive-negative ion pair. Ionic compound: An ionic compound occurs when a negative ion (an atom that has gained an electron) joins with a positive ion (an atom that has lost an electron) general-chemistry.pdf

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